INTRODUCTION

Carbon nanotubes (CNT) discovery by Iijima in 1991. Carbon nanotubes (CNTs) are hollow cylinders of graphite carbon atoms. These tubes are on the nano scale (10-9 m), which is so small that 10,000 of them could fit within the diameter of one human hair. Carbon nanotubes are a new form of carbon with unique electrical and mechanical properties. They can be considered as the result of folding graphite layers into carbon cylinders. Multi walled nanotubes (MWNT) consist of multiple rolled layers (concentric tubes) of grapheme. There are two model, sheets of graphite are arranged in concentric cylinders, e.g., a (0.8) single walled nanotube (SWNT) within a larger (0.17) single-walled nanotube. Recent discoveries of various forms of CNTs have stimulated research on their applications in varied fields. The nature of bonding in carbon nanotubes (CNTs) is described by applied chemistry, specifically, orbital hybridization. This chemical bonding is composed entirely of sp² bonds, similar to those of graphite. This bonding structure, which is stronger than the sp³ bonds found in diamond, provides the molecules with their particular strength. CNTs are the strongest and hard materials on earth, in terms of supple strength and elastic modulus respectively. This strength results from the covalent sp² bonds formed between the individual carbon atoms. As for
thermal conduction, the CNT surpasses even that of diamond, reaching almost double the value of diamond. Carbon nanotubes (CNT) possess many particular characteristics that promise to revolutionize the world of structural materials resulting in significant impact on our capability to build lighter, higher performance structures for aerospace and many other industrial applications. Dexamethasone is a type of steroid drug known as a glucocorticosteroid. It is a man-made version of a natural hormone produced by the adrenal glands. Glucocorticosteroids have a wide range of actions on many parts of the body. The ways in which they cause these many different effects aren't exactly clear. The main effects of dexamethasone and steroids like it seem to be due to their anti-inflammatory properties and their ability to alter immune system responses. For example, dexamethasone helps prevent white blood cells from traveling to areas of the body where they might add to swelling problems (such as around tumors). The reasons for some of it's other actions (such as increasing appetite and reducing nausea) aren't as clear. Designated chemically as 9-Fluoro-11,17-dihydroxy-17-(2-hydroxyacetyl)-10,13,16-trimethyl6,7,8,9,10,11,12,13,14,15,16,17-dodecahydro-3H-cyclopenta[a]phenanthren-3-one. The molecular weight is 392.461 g/mol. Its molecular formula is C_{22}H_{29}FO_{5}. The structural formula is:

![Scheme 1](image)

**EXPERIMENTAL**

**Substances**
Ethanol (purity %96) as solvent, activated carbon and multi-wall carbon nanotube with 95% pure degree, production of neutrino Company. Dexamethasone with 99% pure degree production of Pharmaceutical companies daro pakhsh.

**Devices used**
In this researcher spectrophotometer UV / VIS Jenway 6505 model England, magnetic stirrer (Heidolph, Mr3001 Model), Analytical balances (Sartorius Model), Filter paper (Albet), were used.

**Adsorption experiments**
In order to provide dexamethasone solution, twice distilled water was use. Multi walled carbon nanotube was used as an adsorbent. First 50 ppm of dexamethasone was provided using this sample, some solutions with different concentrations of 4 to 12 mg/lit of pure dexamethasone were prepared specific amount of carbon nanotube (0.005 gr) was added to flasks containing dexamethasone, and dye as an adsorbent. It was stirred, using a magnetic stirrer for 10 minutes. Then liquid and solid phase were separated by means of a filter paper. The concentration of dexamethasone was measured by using spectrophotometer tool adsorption rate, gained for dexamethasone. All tests have been performed at the lab with the temperature of (25 ± 2°C). In a two component system consisting of adsorber solution, a graph of the concentration of dissolved solid phase qe (mg / g) was dissolved in the solution concentration Ce (mg/L) in an equilibrium adsorption isotherm is expressed. In a liquid-solid adsorption from solution on to the surface of the separation of dissolved solids remaining dissolved in the solution until the solution reaches the dynamic equilibrium on a solid surface. The Freundlich, Langmuir and Temkin models are employed to analysis adsorption occurred in the experiment. Adsorption isotherms to describe the adsorption capacity to facilitate the process of evaluating the feasibility of the intended application and is useful for the analysis and design of adsorption systems.

**Langmuir model**
The Langmuir adsorption model is the
most common model used to quantify the amount of adsorbs ate on an adsorbent as a function of partial pressure or concentration at a given temperature. This equation expressed by relation.

\[ \frac{C_e}{q_e} = \frac{1}{q_m} + \frac{1}{q_m b} C_e \]  

(1)

In this equation, \( q_e (\text{mg.g}^{-1}) \) is the solution was adsorbed the surface and \( q_m \) is equilibrium constant of adsorption and \( b \) is the capacity of adsorption in saturated single layer and \( C_e (\text{Mg.} \cdot \text{L}^{-1}) \) Is solution in equilibrium state?

Freundlich model

The Freundlich equation or Freundlich adsorption isotherm is an adsorption isotherm, which a curve is relating the concentration of a solute on the surface of an adsorbent, to the concentration of the solute in the liquid with which it is in contact. In 1909, Freundlich gave an empirical expression representing the isothermal variation of adsorption of a quantity of gas adsorbed by unit mass of solid adsorbent with pressure. This equation is known as Freundlich adsorption isotherm or Freundlich adsorption equation. This model is specified with equation.

\[ q_e = K_n^{1/n} C_e^{1/n} \ln q_e = \ln K_n + \frac{1}{n} \ln C_e \]  

(2)

In this equation, \( q_e (\text{mg.g}^{-1}) \) is amount of absorbed material in absorbent surface, \( K_n \) and \( n \) in arrangement are adsorption capacity and adsorption intensification.

Temkin model

The Temkin model is linearly represented as equation (3) and generally applied in the form:

\[ q_e = B \ln A + B \ln C_e \]  

(3)

Where \( A \) and \( B \) are the Temkin isotherm constant (\text{L/g}) and heat of sorption (\text{J/mol}) respectively. \( R \) is the gas constant (\text{J/mol/k}), \( b \) is the Temkin isotherm constant linked to the energy parameter, \( B \), as shown on equation:

\[ b = RT/B \]  

(4)

\( T \) is the absolute temperature in Kelvin.

RESULTS

Table 1 shown the absorbance of dexamethasone in the absent of activated carbon and carbon nanotube in 238 nm wavelength. As seen, table 1 shown the amount of absorbance is increased when concentrations are increased. Table 2 shown the increasing absorbance of dexamethasone against the concentrations in the presence of carbon nanotube. Table 3 also shown the increasing absorbance of dexamethasone

<table>
<thead>
<tr>
<th>Concentrations (Mg.L^{-1})</th>
<th>Absorbance of Dexamethasone(nm)</th>
</tr>
</thead>
<tbody>
<tr>
<td>8</td>
<td>0.51482</td>
</tr>
<tr>
<td>10</td>
<td>0.61285</td>
</tr>
<tr>
<td>12</td>
<td>0.71986</td>
</tr>
<tr>
<td>14</td>
<td>0.82025</td>
</tr>
</tbody>
</table>

Table 2: The amount of Dexamethasone on 0.005 grams of carbon nanotubes

<table>
<thead>
<tr>
<th>Concentrations (Mg.L^{-1})</th>
<th>Material adsorbed on 0.01 grams of carbon nanotubes(mg.g^{-1})</th>
</tr>
</thead>
<tbody>
<tr>
<td>8</td>
<td>0.31128</td>
</tr>
<tr>
<td>10</td>
<td>0.459911</td>
</tr>
<tr>
<td>12</td>
<td>0.56124</td>
</tr>
<tr>
<td>14</td>
<td>0.66988</td>
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</tbody>
</table>

Table 3: The amount of Dexamethasone on 0.005 grams of activated carbon

<table>
<thead>
<tr>
<th>Concentrations (Mg.L^{-1}(0.01))</th>
<th>Material adsorbed on grams of activated carbon (mg.g^{-1})</th>
</tr>
</thead>
<tbody>
<tr>
<td>8</td>
<td>0.3035</td>
</tr>
<tr>
<td>10</td>
<td>0.4100121</td>
</tr>
<tr>
<td>12</td>
<td>0.51036</td>
</tr>
<tr>
<td>14</td>
<td>0.62251</td>
</tr>
</tbody>
</table>
Table 4: Parameters of Langmuir, Freundlich and Temkin isotherms of the Dexamethasone on Carbon nanotubes and Activated carbon

<table>
<thead>
<tr>
<th></th>
<th>Langmuir</th>
<th>Freundlich</th>
<th>Temkin</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>b</td>
<td>q</td>
<td>R²</td>
</tr>
<tr>
<td>Carbon nanotubes</td>
<td>0.36</td>
<td>1.25</td>
<td>0.994</td>
</tr>
<tr>
<td>Activated carbon</td>
<td>1.5</td>
<td>3.38</td>
<td>0.996</td>
</tr>
</tbody>
</table>

Fig. 1: Absorbance of Dexamethasone in the absence of CNT & AC

Fig. 2: Effect of initial concentration on absorbent CNT & AC Curve adsorption isotherm

Fig. 3: Freundlich isotherm of Dexamethasone on CNT & AC

Fig. 4: Temkin isotherm of Dexamethasone on CNT & AC

Fig. 5: Langmuir isotherm of Dexamethasone on CNT & AC
against the concentrations in the presence of activated carbon. Calculated parameters of these models are shown in Table 4 and the Langmuir, Freundlich and Temkin isotherms of the adsorption process of dexamethasone on AC and CNT are shown in figures 3 to 5. It was observed that the experimental data were well represented by Langmuir Freundlich and Temkin models.

**Study of adsorption isotherm**

Calculated parameters based on exothermic equations are shown in Table 4. As seen, the values of correlation coefficient ($R^2$) for the absorbance of dexamethasone with the use of carbon nanotube for Langmuir, Freundlich and Temkin models are 0.994, 0.92 and 0.896 and with the use of activated carbon are 0.996, 0.82 and 0.8264, respectively. The results showed that the absorbance of dexamethasone on both absorbents is in conformance with Langmuir and Freundlich isotherm.

**CONCLUSION**

In this study, we compare the adsorption isotherms of dexamethasone by activated carbon and carbon nanotube. Based on obtained results we conclude that nanotube has more efficiency in removal of dexamethasone rather than activated carbon. Results of isothermic experiments showed that the correlation coefficient of Langmuir isotherm's equation for carbon nanotube was more than activated carbon. The values of RL in carbon nanotube and activated carbon were less than 1, indicating that the absorbance is appropriate for both adsorbents. Therefore, in total, it is concluded that correlation coefficient (RL) in Langmuir isotherm model for carbon nanotube were higher and its efficiency in the removal of dexamethasone is better than activated carbon.

**REFERENCES**

13. M. Vadi, M. Zakeri. "Thermodynamic study of
Chrysens adsorption by multi-walled on carbon nanotube”. Oriental Journal of Chemistry. 27(3): 945-951


